

colligative properties yahoo answers pdf

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Best Answer: Three colligative properties are boiling point, freezing point, and vapor pressure. They are called colligative properties because they are related to the number and energy of collisions between particles and not what the particles are.

Colligative properties? | Yahoo Answers

Colligative properties are properties of solutions that depend on the number of particles in a given volume of solvent and not on the mass of the particles. Colligative properties include: lowering of vapor pressure; elevation of boiling point; depression of freezing point; osmotic pressure (see Osmosis; Reverse Osmosis).

Colligative properties? | Yahoo Answers

The four commonly studied colligative properties are freezing point depression, boiling point elevation, vapor pressure lowering, and osmotic pressure. Since these properties yield information on the number of solute particles in solution, one can use them to obtain the molecular weight of the solute.

Colligative Properties- Chemistry? | Yahoo Answers

Colligative: particles are particles Colligative comes from colligate "to tie together Colligative properties have common origin Colligative properties depend on amount of solute but do not depend on its chemical identity Solute particles exert their effect merely by being rather than doing The effect is the same for all solutes

Colligative Properties - College of DuPage

Colligative Properties- Page 1 Lecture 4: Colligative Properties By definition a colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent matters but the kind of solute does not matter.

Colligative Properties- Page 1 Lecture 4: Colligative

Colligative Properties Problems - Answers Remember that Colligative properties depend only on the number or concentration of particles in a solution. The properties are, for ideal solutions, independent of the kind or size of the particles, whether ionic, large or small etc.

Colligative Properties Problems - Answers - HSU Users Web

Osmotic pressure is a colligative property that can be used to determine the molar mass of an unknown substance. The osmotic pressure is determined by measuring the height of the column of solution and converting this value to mm of Hg (1 mm Hg = 1 torr, 760 torr = 1 atm). Using, $\hat{\Pi} V \text{ solution} = n \text{ solute} R T$ (where $\hat{\Pi}$ is the osmotic pressure, R

CHEMISTRY 142 " Example Problems

3 Colligative Properties January 13 Colligative Properties It is known that: Dissolved solute in pure liquid will change the physical property of the liquid, i.e., Density, Bpt, Fpt, Vapor Pressure This new developed property is called Colligative Property; a method of counting the number of solute in solution. Vapor Pressure

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